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Answer the following questions with regards to the Heating Curve.

1) What is happening to the average potential energy of the molecules in the sample during section 3? Potential Energy remains constant. It only increases or decreases when

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plateaus. ...

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$\text{SO}_2(\text{g}) + 2 \text{HNO}_3(\text{aq}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow 3 \text{H}_2\text{SO}_4(\text{aq}) + 2 \text{NO}(\text{g})$ a.
Calculate ΔH for the above reaction. Based on the enthalpy change, determine whether the reaction is endothermic or exothermic and whether it would be thermodynamically favored or unfavored.

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block of uranium metal
(specific heat = $0.117 \text{ J g}^{-1} \text{ } ^\circ\text{C}^{-1}$) at $200.0 \text{ } ^\circ\text{C}$
is dropped into 1.00 L

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of deuterium oxide
(also called "heavy
water") (specific heat
 $= 4.211 \text{ J g}^{-1} \text{ } ^\circ\text{C}^{-1}$)
at $25.5 \text{ } ^\circ\text{C}$.

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Name _____ MULTIPLE

CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure.

A) endothermic,

positive

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True or False If the
answer is false, you
must explain why. 1.

_____ Chemical energy
is a form of kinetic

energy. 2. _____ When
we mix water at 100 C
with water at 20 C the
final water temp. is
always 60 C. 3. _____

When we put a 100g

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piece of metal at 100 C
into 100g of water at
20 C the final temp is
20 C. 4.

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The _____ is the set of properties that describe the system. state. The _____ is comprised of all the matter in the reaction. system

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the correct statement
below. a. calorimeter d.
enthalpy b. calorie e.
specific heat c. joule f.
heat capacity ____ 1.
quantity of heat
needed to raise the
temperature of 1 g of
water by 1 C 2. ____ SI
unit of energy
3. ____ quantity of heat
needed to change the
temperature of 1 g of a
...

Chapter 6 Thermochemistry

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Thermochemistry
Exam1 and Problem
Solutions 1. Which
ones of the following
reactions are
endothermic in other
words ΔH is positive? I.
 $\text{H}_2\text{O}(\text{l}) + 10,5\text{kcal} \rightarrow$
 $\text{H}_2\text{O}(\text{g}) \Delta H1$ II. 2NH_3
 $+22\text{kcal}$

Thermochemistry
Exam1 and Problem
Solutions | Online ...
Thermochemistry Quiz
Review Name: _____ 1)

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How many joules are required to raise 20.0 g of ice from $-10.0\text{ }^{\circ}\text{C}$ to water at 50.0 ° ?

(Answer: 11274 J) 2)

How much heat is required to raise the...

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5.2: The First Law of Thermodynamics. First law of thermodynamics: aka

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The Law of Conservation of Energy; a statement of our experience that energy is conserved in any process. We can express the first law in many ways. One of the more useful expressions is that the change in internal energy, ΔE , of a system in any process is equal to the heat, q , added to the system, plus the work, w , done ...

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5.S:

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LO MAPPED – The FRQ's that are listed on a worksheet marked as 'LO MAPPED', are mapped to SPECIFIC LO's in the May 2019 CED. It should be noted that NONE of the questions prior to 2020 were actually mapped by the College Board to the May 2019 CED,

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since they were all written prior to that document being produced.

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