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CHAPTER 15 REVIEW Acid-Base Titration and pH SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Below is a pH curve from an acid-base titration. On it are labeled three points: X, Y, and Z. a. Which point represents the equivalence point? b. At which point is there excess acid in the system? c.

15 Acid-Base Titration and pH

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Modern Chemistry Chapter 15 Acid-Base Titration & pH Section 1 self-ionization of water occurs when two water molecules produce a hydronium (H_3O^+) and a hydroxide (OH^-) ion $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$. The ionization constant (K_w) of water is: $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ M}$ Acidic, Basic, & Neutral IF $[\text{H}_3\text{O}^+] > [\text{OH}^-]$ then solution is acidic.

Modern Chemistry Chapter 15 Acid-Base Titration & pH

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Modern Chemistry Interactive Reader ChapTEr 1 Matter and Change 2 1 2 3 Chemistry Is a Physical Science 3 ... Chapter Review 469 ChapTEr 15 Acid-Base Titration and pH 472 1 2 ... 2 CHAPTER 1. SECTION 1.1 Physical science is the study of nonliving objects and materials. Physics, astronomy, and geology are examples of

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A nitric acid solution is found to have a pH of 2.70 ...

Chemistry Principles of Modern Chemistry When ozone (O_3) is placed in contact with dry, powdered KOH at $-15^\circ C$, the red-brown solid potassium ozonide (KO_3) forms, according to the following balanced equation: $5 O_3 (g) + 2 KOH (s) \rightarrow 2 KO_3 (s) + 5 O_2 (g) + H_2 O (g)$
Calculate the volume of ozone needed (at a pressure of 0.134 atm and $-15^\circ C$) to produce 4.69 g KO_3 .

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