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Stoichiometry Practice Test Short

Answer: Aluminum bromide can be prepared by the reaction of aluminum metal with bromine gas shown by the equation: $2 \text{Al} + 3 \text{Br}_2 \rightarrow 2 \text{AlBr}_3$ Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide that can be produced from 5.6 mol of Al. 2.

Stoichiometry Practice Test

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Extra Stoichiometry Problems 1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.
a. Write and balance the chemical equation. $2 \text{AgNO}_3 + \text{BaCl}_2 \rightarrow 2 \text{AgCl} + \text{Ba}(\text{NO}_3)_2$
b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many

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Modern Chemistry 88 Chapter Test Chapter: States of Matter ... PART I In the space provided, write the letter of the term or phrase that best completes

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each statement or best answers each question. ____ 1. According to the kinetic-molecular theory, particles of an ideal gas ... Stoichiometry, pp. 74–83
TEST A 1. a 2. b 3. a 4. 5. d 6. c 7. a ...

Assessment Chapter Test B - Ed W. Clark High School

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Stoichiometry Chapter Test

Stoichiometry SECTION 2 PROBLEMS

Write the answer on the line to the left.

Show all your work in the space

provided. 1. 4.5 mol The following

equation represents a laboratory

preparation for oxygen gas: $2\text{KClO}_3(\text{s})$

$\rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O

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2 form if 3.0 mol of KClO_3 are totally consumed?

mc06se cFMsr i-vi

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